

Size-Dependent Optical Properties of Colloidal PbS Quantum Dots

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For more than 25 years, interest in colloidal semiconductor nanocrystals or quantum dots (Qdots) has kept growing at an increasing pace, mostly driven by the exceptional optical properties these Qdots possess. Several materials can now be produced by well-controlled colloidal syntheses,¹ and their basic material properties^{2,3} (most importantly the particle size and its influence on the band gap) have been thoroughly investigated. Consequently, more and more application-oriented research is being pursued, primarily in the direction of Qdot-based biolabeling,^{4,5} LEDs,⁶ lasers,^{7,8} and solar cells.⁹ It is expected that these applications will benefit greatly from the enhanced and tunable Qdot optical properties.

In order to fabricate and optimize these devices, it is highly desirable that one has access to the concentration c_0 of the suspended Qdots. Through Beer's law, c_0 can be conveniently determined from the absorbance spectrum of a suspension if the Qdot molar extinction coefficient ϵ is known. At present, ϵ values for cadmium chalcogenide,^{10–12} InP,¹³ InAs,¹⁴ and lead chalcogenide^{15–17} Qdots have been reported. Unfortunately, the data also show that the methods used to determine ϵ still vary strongly among research groups, resulting for instance in different empirical size dependencies of ϵ at the band gap for the same material.^{10,12,16,17}

Experimental methods to determine ϵ and c_0 are separable into two main groups. Often, the value of ϵ at the band gap is used,^{12,15,17} either at a single wavelength or integrated over the first absorption peak. On the other hand, several groups have recently shown that the value of ϵ at energies far above the band gap bears more practical

ABSTRACT We quantitatively investigate the size-dependent optical properties of colloidal PbS nanocrystals or quantum dots (Qdots), by combining the Qdot absorbance spectra with detailed elemental analysis of the Qdot suspensions. At high energies, the molar extinction coefficient ϵ increases with the Qdot volume d^3 and agrees with theoretical calculations using the Maxwell–Garnett effective medium theory and bulk values for the Qdot dielectric function. This demonstrates that quantum confinement has no influence on ϵ in this spectral range, and it provides an accurate method to calculate the Qdot concentration. Around the band gap, ϵ only increases with $d^{1.3}$, and values are comparable to the ϵ of PbSe Qdots. The data are related to the oscillator strength f_{if} of the band gap transition and results agree well with theoretical tight-binding calculations, predicting a linear dependence of f_{if} on d . For both PbS and PbSe Qdots, the exciton lifetime τ is calculated from f_{if} . We find values ranging between 1 and 3 μ s, in agreement with experimental literature data from time-resolved luminescence spectroscopy. Our results provide a thorough general framework to calculate and understand the optical properties of suspended colloidal quantum dots. Most importantly, it highlights the significance of the local field factor in these systems.

KEYWORDS: semiconductor nanocrystals · lead chalcogenide · PbSe · molar extinction coefficient · oscillator strength · exciton lifetime

significance. More specifically, for CdSe,¹¹ InAs,¹⁴ and PbSe¹⁶ Qdots, bulk-like optical properties were measured at 350, 450, and 400 nm, respectively. These were modeled by the Maxwell–Garnett (MG) effective medium theory,^{18,19} applicable to dispersed particles in a (transparent) dielectric host. Using the dielectric function of the respective bulk semiconductor and the solvent refractive index, the MG model yielded an ϵ , or equivalently an absorption coefficient μ or cross section C_{abs} , which agrees well with the experimental data at short wavelengths. This demonstrates that quantum confinement has no influence on ϵ in this spectral range. Hence, ϵ is not affected by the particle size dispersion, improving the accuracy of the reported data.

The successful application of the MG effective medium theory also highlights that it can give us a better quantitative insight in the Qdot optical properties and their

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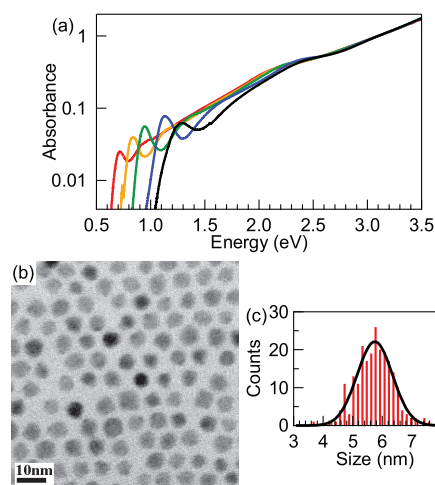


Figure 1. (a) Series of absorbance spectra of PbS Qdots suspended in C_2Cl_4 , normalized at 3.1 eV. All spectra coincide at high energies. (b) Overview TEM image of a typical PbS Qdot suspension. (c) Corresponding size histogram. The mean size equals 5.7 nm, with a size dispersion of 10%.

evolution with the particle size d .²⁰ Indeed, for PbSe Qdots, we recently reported that, in contrast with the bulk-like μ at high energies, the energy-integrated μ at the band gap shows a strong size dependence, resulting in an oscillator strength f_{if} per particle, which increases approximately linearly with d .¹⁶ Experimental data are in agreement with theoretical tight-binding calculations,²¹ and they for instance demonstrate that, for Qdots in the strong confinement regime, a constant f_{if} should not be expected.^{22,23}

In this paper, we apply the methods derived previously¹⁶ to calculate the optical properties of colloidal PbS Qdots. First, the PbS Qdot concentration c_0 is carefully determined by elemental analysis of the Pb concentration of digested PbS Qdot suspensions, in combination with a measurement of the atomic Pb/S ratio. Beer's law then yields the PbS Qdot ε , which is analyzed both at high energies and at energies around the band gap. We make a detailed comparison with the ε of PbSe Qdots. Practically, this allows one to determine which material, from an optical perspective, might be more suitable for further applications. Theoretically, it improves the insight in the underlying physics determining the optical properties, and it confirms the applicability of the MG effective medium theory. Finally, we expand the current analysis methods to obtain the exciton lifetime τ for both materials and compare these calculated values to experimental literature data.

RESULTS

Determination of the Sizing Curve. PbS is a narrow gap semiconductor, with a bulk band gap of 0.41 eV, and a large exciton Bohr radius (18 nm). Consequently, the PbS Qdot band gap E_0 can be tuned over the entire NIR spectral range. Figure 1a shows a series of absorbance spectra for PbS Qdots, normalized at 3.1 eV (400 nm). It demonstrates that the synthesis allows the prepara-

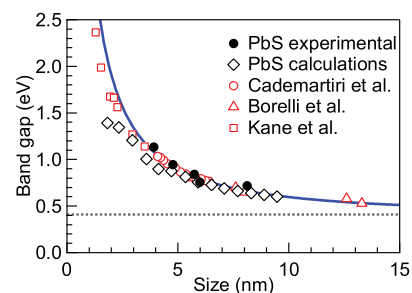


Figure 2. Relation between the PbS Qdot band gap and the particle size. The experimental data (●) agree well with literature values (○, △) and tight-binding calculations (our data ◇, literature data □). To the set of experimental data, a sizing curve is fitted (full line). The dotted line denotes the bulk PbS band gap (0.41 eV).

tion of particles with E_0 varying between 0.71 and 1.28 eV (1740–970 nm). A typical transmission electron microscopy (TEM) image (Figure 1b) illustrates that the PbS Qdots have a spherical shape. An analysis of the size (Figure 1c) yields $d = 5.7$ nm, with a relative size dispersion $\sigma_d = 10\%$ for the particle suspension shown in Figure 1b.

The value of d is measured for five samples and correlated with E_0 to construct a sizing curve. Figure 2 shows that the data correspond well with those of Cademartiri *et al.*¹⁵ and Borrelli and Smith.²⁴ A fit to both our data and the experimental literature values yields the following result (size range fitted: $d = 3.9$ –13.3 nm):

$$E_0 = 0.41 + \frac{1}{0.0252d^2 + 0.283d} \quad (1)$$

The construction of this sizing curve allows us to determine d directly from E_0 , avoiding a lengthy TEM analysis for each sample synthesized. Experimental results agree well with theoretical tight-binding calculations (Figure 2). Both predict an E_0 scaling mainly with d^{-1} . This is similar to the size dependence of E_0 for PbSe Qdots,^{16,21} albeit that PbS Qdots are typically larger than PbSe Qdots for a given E_0 . Note that, although we have not synthesized particles with a size below 3.5 nm, the extrapolation of the sizing curve still agrees well with calculated data of Kane *et al.*²⁵

Assessment of the Sample Purity. When determining the Qdot concentration using elemental analysis of the cation and anion, it is crucial that samples are free of unreacted precursors. As described by Cademartiri *et al.*,¹⁵ lead chloride ($PbCl_2$) is insoluble in apolar solvents, and trace amounts still present after synthesis and purification precipitate slowly from the suspension. In accordance with their procedure, we use powder X-ray diffraction (XRD) to demonstrate the absence of $PbCl_2$. A typical X-ray diffractogram (XRD) is shown in Figure 3a. It confirms that all $PbCl_2$ is removed, as its presence would be revealed by, for instance, reflections at 22.9, 35.8, and 46.5° 2θ .²⁶ The position of the reflections aris-

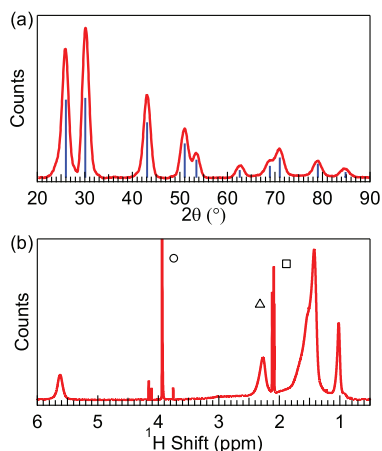


Figure 3. (a) X-ray diffractogram of PbS Qdots. The PbS Qdot crystal structure corresponds to the bulk PbS rocksalt structure (vertical lines). In addition, no reflections due to unreacted PbCl₂ are observed. (b) ¹H NMR spectrum of a PbS Qdot suspension in toluene-*d*₈. Apart from CH₂Br₂ (○), residual toluene (□) and possible methanol traces (△), no sharp resonances arising from unreacted OLA-S and/or TOPS are observed. The broad resonances correspond to the OA ligands.

ing from the PbS Qdots agrees well with the expected values for bulk PbS (rocksalt structure, lattice constant $a = 5.936 \text{ \AA}$),²⁶ demonstrating that these colloidal Qdots maintain their crystal structure when the size is decreased from bulk down to the nanoscale.

To confirm that the suspensions are free of sulfur-based precursors (oleylamine–sulfur complexes, OLA-S, and/or tri-*n*-octylphosphine sulfide, TOPS), PbS Qdots are suspended in deuterated toluene (toluene-*d*₈) and a quantitative proton nuclear magnetic resonance (¹H NMR) spectrum is measured.^{16,27} Figure 3b shows that, apart from a sharp singlet due to dibromomethane (CH₂Br₂, 3.94 ppm, ○), a quintuplet due to residual toluene (2.09 ppm, □) and a singlet possibly due to methanol (2.12 ppm, △), no sharp signals arising from OLA-S and/or TOPS are observed. The broad resonances in fact arise from oleic acid (OA) ligands bound to the PbS Qdot surface. A similar ¹H NMR spectrum is observed for OA-capped PbSe Qdots.²⁷

Both the XRD and NMR results clearly demonstrate that our samples are free of unreacted precursors. This enables us to reliably determine PbS Qdot concentrations using elemental analysis.

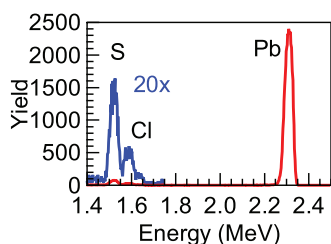


Figure 4. RBS spectrum of a typical PbS Qdot close-packed thin film. The three peaks can be attributed to He⁺ ions that have collided with Pb, S, and Cl. The S and Cl peaks have been multiplied by 20 for clarity.

TABLE 1. Pb/S, Cl/Pb, and Cl/Pb_e Ratio Obtained for the Five Samples Used for RBS (Pb_e Denotes the Excess Pb Atoms)

<i>d</i> (nm)	Pb/S	Cl/Pb	Cl/Pb _e
3.7	1.37	0.49	1.80
4.1	1.23	0.38	2.07
4.9	1.26	0.37	1.83
5.9	1.27	0.32	1.48
6.8	1.28	0.34	1.55

Qdot Composition and Concentration. We use Rutherford backscattering spectrometry (RBS)^{28,29} to measure the Pb/S atomic ratio R of five samples. Figure 4 shows a typical RBS spectrum, measured on 5.9 nm PbS Qdots. Three peaks are observed, corresponding to He⁺ ions, which have collided with Pb, S, and Cl. From the area under the peaks, we determine the respective atomic ratios. Table 1 summarizes the results. We estimate the uncertainty on the data to be 5%. The sample with the smallest size shows a Pb/S ratio of 1.37:1. For all other samples, we obtain, within experimental error, a constant ratio of 1.26:1. These results demonstrate that, like PbSe Qdots,¹⁶ PbS Qdots are nonstoichiometric, showing a systematic Pb excess Pb_e. Interestingly, a significant amount of Cl is also present despite the absence of unreacted PbCl₂, at ratios which are consistent with results of Cademartiri *et al.*³⁰ From the approximate 2:1 ratio of Cl to Pb_e, we conclude that the excess Pb atoms most probably reside at the Qdot surface, with the Cl atoms bound to them.

Assuming a spherical particle, the number of atoms per Qdot N equals

$$N = \frac{4\pi}{3} \left(\frac{d}{a} \right)^3 \quad (2)$$

Taking the Pb/S stoichiometry R into account, the Pb weight concentration C_{Pb} (mg/L), determined with inductively coupled plasma mass spectrometry (ICP-MS), and the Pb molar mass $M_{\text{Pb}} = 207.2 \text{ g/mol}$ then yield the Qdot concentration c_0 (μmol/L):

$$c_0 = \frac{10^3}{N} \frac{1 + R}{R} \frac{C_{\text{Pb}}}{M_{\text{Pb}}} \quad (3)$$

Table 2 shows the results for the four samples measured. Note that, in accordance with the RBS measurements, we take $R = 1.37:1$ for the 3.6 nm PbS Qdots,

TABLE 2. Weight Concentration C_{Pb} , Obtained with ICP-MS, Pb/S Ratio R (In Accordance with the RBS Data), Resulting Qdot Concentration c_0 , and Absorbance of the Suspension at 400 nm A_{400}

<i>d</i> (nm)	σ_d (%)	C_{Pb} (mg/L)	R	c_0 (μM)	A_{400} (cm ⁻¹)
3.6	9.6	5.78	1.37	0.054	0.563
4.5	10.2	7.08	1.26	0.033	0.707
5.4	8.8	7.65	1.26	0.021	0.761
6.5	10.7	4.98	1.26	0.008	0.504

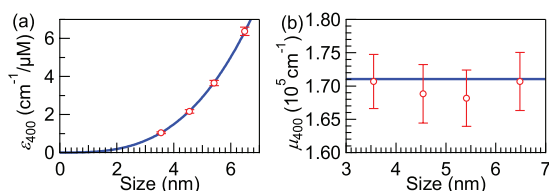


Figure 5. (a) At 400 nm, the PbS Qdot molar extinction coefficient ϵ_{400} increases with d^3 . (b) Consequently, the absorption coefficient μ_{400} is size-independent. Results are in agreement with a theoretical μ , calculated using the bulk PbS dielectric function (full line).

while for the other samples, $R = 1.26:1$ is used. σ_d is listed as well to highlight that the size-dependent optical properties are calculated starting from samples with a small size dispersion.

PbS Qdot Molar Extinction Coefficient at 400 nm. From the absorbance of an equal amount of PbS Qdots, the molar extinction coefficient is derived using Beer's law. The absorbance at 400 nm A_{400} is listed in Table 2. At this wavelength, the molar extinction coefficient scales with the nanocrystal volume (Figure 5a):

$$\epsilon_{400} = (0.0233 \pm 0.0001)d^3 \text{ cm}^{-1}/\mu\text{M} \quad (4)$$

with d in nanometers. The following relation relates the experimental ϵ_{400} to the PbS Qdot absorption coefficient μ_{400} (N_A : Avogadro's constant):

$$\mu_{400} = \frac{6 \ln(10)\epsilon_{400}}{\pi d^3 N_A} \quad (5)$$

Figure 5b shows the resulting size-independent μ_{400} for PbS Qdots. The uncertainty on the data is estimated to be 3%.

Using the MG effective medium theory, μ can also be calculated at a given wavelength λ from the dielectric function of the semiconductor ($\epsilon_R + i\epsilon_I$) and the solvent ($\epsilon_s = n_s^2$):

$$\mu = \frac{2\pi}{n_s \lambda} |f_{LF}|^2 \epsilon_I \quad (6)$$

$$f_{LF} = \frac{3\epsilon_s}{\epsilon_R + i\epsilon_I + 2\epsilon_s} \quad (7)$$

Here, f_{LF} denotes the (complex) local field factor. Equation 6 shows that it can strongly modify the absorption of a material when it is dispersed as small particles in a host with a different dielectric function (this is a consequence of dielectric confinement).^{31,32} Using bulk values for the PbS dielectric function at 400 nm ($\epsilon_R = 4.53$, $\epsilon_I = 26.4$)³³ and the tetrachloroethylene (C_2Cl_4) refractive index ($n_s = 1.53$),³⁴ eq 6 yields a size-independent $\mu_{400} = 1.71 \times 10^5 \text{ cm}^{-1}$ (Figure 5b, full line). Using the theoretical μ_{400} , we find the following theoretical expression for the molar extinction coefficient: $\epsilon_{400} = 0.0234d^3 \text{ cm}^{-1}/\mu\text{M}$. Clearly, both μ_{400} and ϵ_{400} are in excellent agreement with the experimental data.

As for PbSe Qdots,¹⁶ the cubic size dependence of ϵ_{400} demonstrates that the PbS Qdot optical properties at energies well above the band gap are not influenced by quantum confinement. This was already apparent from the normalized absorbance spectra, as all spectra coincide at high energies (Figure 1). These results have an important practical consequence. As they are independent of size and therefore size dispersion, ϵ_{400} allows for a very accurate determination of the Qdot concentration c_0 . This method is therefore to be preferred over a procedure based on the Qdot absorbance at the band gap.^{12,17}

PbS Qdot Optical Properties at the Band Gap. We calculate the molar extinction at the band gap ϵ_{gap} by integrating the first absorption peak on an energy scale.¹⁶ The four ICP-MS samples directly yield ϵ_{gap} from an integration of the ϵ spectrum. However, we have measured the absorbance spectrum of 10 additional PbS Qdot samples suspended in C_2Cl_4 , and for these samples, ϵ_{gap} can be determined using ϵ_{400} (A_{gap} is the area of the energy-integrated first absorption peak):

$$\epsilon_{\text{gap}} = \frac{A_{\text{gap}}}{A_{400}} \epsilon_{400} \quad (8)$$

The results are shown in Figure 6a. Note that the data are more scattered than the values at 400 nm, confirming that a determination of c_0 using Beer's law and ϵ at 400 nm will be more accurate.

In contrast with ϵ_{gap} , which depends on n_s through f_{LF} , the Qdot oscillator strength f_{if} is a material property. It therefore provides a better insight into the intrinsic Qdot optical properties. Approximating the Qdot as

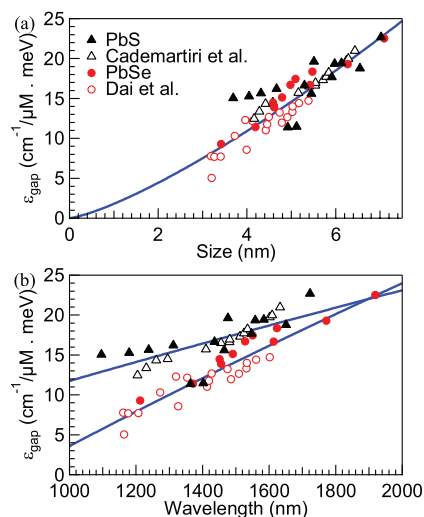


Figure 6. (a) At the band gap, we observe a similar molar extinction coefficient ϵ_{gap} for PbS and PbSe Qdots of equal size. Both scale with $d^{1.3}$. The results are in agreement with PbS data of Cademartiri *et al.* (after correction for their observed Pb/S Qdot ratio) and PbSe data of Dai *et al.* (b) When comparing PbS and PbSe Qdots of equal band gap, PbS Qdots clearly have a larger ϵ_{gap} . This is due to their different sizing curve, resulting in larger PbS Qdots for a given band gap.

a two-level system around the band gap, it can be calculated directly from ε_{gap} (m_e : free electron mass, ε_0 : electric constant):¹⁶

$$f_{\text{if}} = \frac{2\varepsilon_0 c m_e n_s}{\pi N_A \hbar e} \frac{1}{|f_{\text{LF}}|^2} \ln(10) \varepsilon_{\text{gap}} \quad (9)$$

As PbS Qdots are suspended in C_2Cl_4 , we use $n_s = 1.49$,³⁴ valid at NIR wavelengths. In accordance with our recently reported calculations for PbSe Qdots, we estimate $|f_{\text{LF}}|^2$ using bulk data for PbS, valid around the PbS band gap: $\varepsilon_{\text{R}} = 20.25 \gg \varepsilon_{\text{I}}$.³³ This yields $|f_{\text{LF}}|^2 = 0.0725$. Figure 7 shows the resulting f_{if} as a function of d . An approximately linear dependence is observed, and results agree well with theoretical tight-binding calculations.

DISCUSSION

Comparison of ε_{gap} with PbSe Qdots and Literature Data. Figure 6a compares the PbS Qdot data to ε_{gap} of Cademartiri *et al.*¹⁵ We have slightly adjusted their values to account for two differences. First, although they observed that PbS Qdots are nonstoichiometric in an earlier report,³⁰ it appears that this was not considered in their more recent calculation of ε_{gap} (data were based solely on a Pb elemental analysis¹⁵). Taking their reported Pb/S ratio of 1.3–1.5, it is estimated that their ε_{gap} is *ca.* 15–20% too small. We therefore choose to increase their values by 17.5%. Second, they report a wavelength-integrated value $\varepsilon_{\text{gap},\lambda}$. This can be converted to an energy-integrated ε_{gap} .¹⁶

$$\varepsilon_{\text{gap}} = \frac{eE_0^2}{\hbar c} \varepsilon_{\text{gap},\lambda} \quad (10)$$

After conversion, we see that both data sets agree well, confirming the validity of our method to calculate ε_{gap} .

For further comparison, we suspended 11 PbSe Qdot samples in C_2Cl_4 and calculated ε_{gap} from eq 8 using $\varepsilon_{400} = 0.0311d^3 \text{ cm}^{-1}/\mu\text{M}$ ($\mu_{400} = 2.27 \times 10^5 \text{ cm}^{-1}$). These equations are obtained using the MG effective medium theory and bulk PbSe data at 400 nm

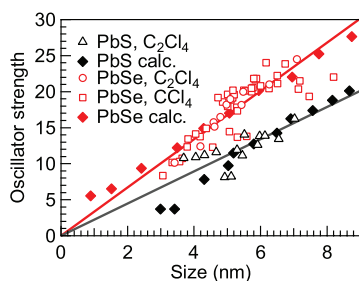


Figure 7. The experimental PbS and PbSe Qdot oscillator strength increases approximately linearly with size. PbSe Qdots suspended in CCl_4 (\square) yield the same values as PbSe Qdots suspended in C_2Cl_4 (\circ). Data for PbS are, however, 33% smaller (\triangle). For both materials, experimental values agree with tight-binding calculations (\blacklozenge).

($\varepsilon_{\text{R}} = -10.67$, $\varepsilon_{\text{I}} = 20.54$),³⁵ as optical properties are also bulk-like at this wavelength. In addition to our data, Dai *et al.*¹⁷ recently reported a wavelength-integrated $\varepsilon_{\text{gap},\lambda}$ (normalized by a standard peak width of $K = 60 \text{ nm}$). Again, this is readily converted to an energy-integrated ε_{gap} through eq 10. After conversion, both data sets agree well. Interestingly, Figure 6a shows that PbS and PbSe Qdots of equal size have a comparable ε_{gap} . A power law fit to both the PbS and PbSe Qdot data yields

$$\varepsilon_{\text{gap}} = 1.8d^{1.3} \text{ cm}^{-1} \text{ meV}/\mu\text{M} \quad (11)$$

From a practical perspective, ε_{gap} at a given wavelength is often more relevant. In Figure 6b, we therefore compare PbS and PbSe Qdots of equal E_0 . It shows that, due to the different size dependence of E_0 , PbS Qdots, which are larger for a given E_0 , also have the larger ε_{gap} . This implies that, for optical applications using colloidal lead chalcogenide Qdots, PbS might be a more suitable material than PbSe.

Comparison of f_{if} with the PbSe Qdot f_{if} . For PbSe Qdots suspended in C_2Cl_4 , we estimate $|f_{\text{LF}}|^2$ at the band gap using $\varepsilon_{\text{R}} = 27.04$,³⁵ yielding $|f_{\text{LF}}|^2 = 0.0446$. This value is 38% smaller than f_{LF} for PbS Qdots, due to the larger ε_{R} of PbSe. Figure 7 compares f_{if} for PbS and PbSe Qdots. First, for PbSe Qdots, we indeed observe that samples suspended in CCl_4 (data taken from our previous report¹⁶) and in C_2Cl_4 show the same f_{if} , confirming that f_{if} is a solvent-independent material property. Second, comparing PbSe to PbS values, we observe that f_{if} is significantly smaller (*ca.* 33%) for PbS Qdots (note that this decreases in f_{if} is, however, almost exactly compensated by the difference in f_{LF} ; this is why a similar ε_{gap} is still observed in Figure 6a). Both data sets agree well with the tight-binding calculations, and we observe that the PbS and PbSe oscillator strength scales approximately linearly with d .

$$f_{\text{if,PbS}} = (2.23 \pm 0.09)d \quad (12)$$

$$f_{\text{if,PbSe}} = (3.35 \pm 0.05)d \quad (13)$$

It shows that a μ -independent f_{if} should not be expected for Qdots in the strong confinement regime, as predicted by earlier calculations within an effective mass approximation.^{22,23}

Calculation of the Exciton Lifetime. The oscillator strength is related to the exciton decay rate τ^{-1} by Fermi's golden rule ($\omega = E_0/\hbar$ is the band gap transition frequency):³⁶

$$\tau^{-1} = \frac{e^2}{2\pi\varepsilon_0 c^3 m_e} n_s |f_{\text{LF}}|^2 \omega^2 \frac{f_{\text{if}}}{g} \quad (14)$$

Here, g denotes the degeneracy of the band gap exciton ($g = 64$ for lead chalcogenide Qdots). We hereby assume that the oscillator strength of spontaneous emission is equal to the oscillator strength of the band gap absorption f_{if} , reduced by g .

Using eq 9, we can determine τ^{-1} directly from ε_{gap} , hereby avoiding the need to calculate f_{if} :

$$\tau^{-1} = \frac{e}{\pi^2 N_A \hbar c^2} n_s^2 \omega^2 \frac{\ln(10) \varepsilon_{\text{gap}}}{g} \quad (15)$$

Figure 8a shows that, with the exception of the smallest PbSe Qdots ($d < 5$ nm), τ^{-1} increases linearly with ω . This linear trend is in agreement with experimental data obtained on CdSe and CdTe Qdots and, from a theoretical viewpoint, is explained by a size-independent dipole matrix element:³⁷

$$|\langle f | \mathbf{p} | i \rangle|^2 = \frac{3 \hbar m_e \omega}{2} f_{\text{if}} \quad (16)$$

The deviation from linearity in the case of small PbSe Qdots therefore indicates that $|\langle f | \mathbf{p} | i \rangle|^2$ becomes size-dependent for $d < 5$ nm.

The exciton lifetime τ is shown in Figure 8b as a function of d . For PbS Qdots suspended in C_2Cl_4 , τ varies linearly between 1 and 1.8 μs over the entire range of sizes that we have analyzed. For PbSe Qdots in CCl_4 , τ varies between 2.0 and 3.4 μs , while somewhat smaller values are obtained when they are suspended in C_2Cl_4 , due to a larger local field factor in C_2Cl_4 . Interestingly, the PbSe Qdot τ shows a minimal value for a diameter of about 5 nm. These data are in reasonable agreement with experimental values obtained through time-resolved luminescence measurements for both PbS³⁸ and PbSe Qdots.^{39,40} More specifically, the decreasing trend in τ for small PbSe Qdots has also been recently observed by Kigel *et al.*,³⁹ who measured radiative lifetimes between 4.97 and 2.74 μs for PbSe Qdots of $d = 2.7\text{--}4.7$ nm,

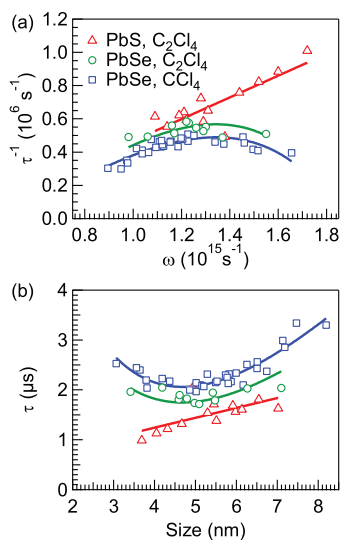


Figure 8. (a) For both PbS (Δ) and PbSe Qdots suspended in C_2Cl_4 (\circ) and CCl_4 (\square), the decay rate τ^{-1} increases linearly with the band gap frequency ω (except for the smallest PbSe Qdots, $d < 5$ nm). (b) The resulting lifetime τ for PbSe Qdots varies between 1.7 and 3.4 μs . Values for PbS Qdots are somewhat smaller (1–1.8 μs).

and Oron *et al.*,⁴⁰ who find 2.5–0.9 μs for the fast component of their luminescence decay ($d = 3.2\text{--}4.3$ nm).

CONCLUSIONS

PbS Qdots are synthesized according to literature methods. Their size is determined with TEM and related to the Qdot band gap to construct a sizing curve. With RBS, we show that the Qdots are nonstoichiometric, with an excess of Pb atoms most probably located at the Qdot surface. A layer of Cl atoms is bound to these excess Pb atoms. Combination of the Pb/S ratio with elemental ICP-MS analysis of the atomic Pb concentration enables us to calculate the Qdot concentration. Through Beer's law, we then obtain the Qdot molar extinction coefficient ε .

At energies far above the band gap, ε scales with the Qdot volume. Its value is compared to a theoretical calculation using the MG effective medium theory and the bulk semiconductor dielectric function. Both show excellent agreement. This has a large practical advantage: Qdot concentrations can be accurately determined from ε in this spectral range, irrespective of size dispersion.

Around the band gap, the energy-integrated ε_{gap} scales with $d^{1.3}$. The PbS Qdot ε_{gap} is clearly larger than ε_{gap} for PbSe Qdots of equal band gap, rendering PbS Qdots more suitable for optical applications due to their stronger absorption at the band gap. The Qdot oscillator strength at the band gap f_{if} can be calculated directly from ε_{gap} . We observe that the PbS Qdot f_{if} is 33% smaller than f_{if} for PbSe Qdots, in agreement with tight-binding calculations. Both scale approximately linearly with size.

From f_{if} , or equivalently ε_{gap} , the exciton lifetime τ is determined. We find values of 1.7–3.4 μs for PbSe Qdots and 1–1.8 μs for PbS Qdots, in reasonable agreement with recent literature data. Interestingly, in contrast with PbS Qdots, small PbSe Qdots show a decrease in lifetime with increasing size, a trend which has recently also been observed using time-resolved photoluminescence spectroscopy. Although experimental values still differ by a factor of 2–3, they clearly show a trend which is in line with our results. In this respect, these calculated lifetimes provide a useful reference in the interpretation of experimental data.

In conclusion, we present a thorough general framework to understand and quantify the optical properties of colloidal semiconductor Qdots, both at high energies and at the band gap. Most importantly, the correspondence with the MG effective medium theory highlights the importance of the local field factor and the effects of dielectric confinement in the determination of ε , f_{if} , and τ .

METHODS

Materials. C_2Cl_4 ($\geq 99\%$), CH_2Br_2 (99+%), and OA (90%) were purchased from Sigma-Aldrich. TOP (97%) was purchased from STREM Chemicals and OLA (80–90%) from Acros. Sulfur (99.999%, S) and $PbCl_2$ (99.999%) were purchased from Alfa-Aesar. Methanol (MeOH), butanol (BuOH), and toluene were all of quality “for synthesis”; concentrated (65%) nitric acid (HNO_3) was of quality “normatom ultrapure”. These products were ordered from VWR. Toluene- d_8 (99.96% deuterated) was purchased from CortecNet.

Colloidal PbS Qdot Synthesis. The PbS Qdot synthesis is based on the procedure described by Cademartiri *et al.*³⁰ The entire synthesis is carried out under a nitrogen atmosphere to avoid oxidation. First, a stock solution of 0.16 g of S, dissolved in 15 mL of OLA, is heated to 120 °C for 30 min and afterward cooled to room temperature. In a three-neck flask, 0.56 g of $PbCl_2$ is mixed with 10 mL of OLA and heated to 75–150 °C (depending on the desired particle size). After 30 min, 3 mL of the S stock solution is mixed with 3 mL of OLA and injected into the three-neck flask. To increase the available size range, 225 μ L of TOP may be added to the synthesis. When the desired growth time is reached (typically 1–20 min), the reaction is quenched by addition of 20 mL of BuOH and 10 mL of MeOH. After centrifugation and decantation, the PbS Qdots are suspended in toluene.

After synthesis, the PbS Qdot organic ligand shell, consisting of OLA, is substituted by OA, by adding 400 μ L of OA to the Qdot suspension, followed by precipitation with an excess of MeOH and centrifugation. After decantation, the Qdots are re-suspended in toluene and precipitated again with MeOH to further remove impurities. After this second centrifugation and decantation step, the OA-capped Qdots are finally suspended in toluene and stored under nitrogen and in the dark.

PbSe Qdots are synthesized based on the procedure developed by Murray *et al.*⁴¹ The synthesis is described in detail elsewhere.¹⁶

Determination of the PbS Qdot Size. We determine the mean diameter d and relative standard deviation σ_d (also called size dispersion) of five PbS Qdot suspensions with TEM. The samples are prepared by dipping a TEM grid in a diluted Qdot suspension. We use multiple images to measure the area of 200–400 particles, from which the equivalent circular diameter is calculated. Values of d and σ_d are calculated directly from the individual sizes but agree well with the results of a Gaussian fit to the histogram.

Purity of the PbS Qdot Suspensions. In accordance with the procedure of Cademartiri *et al.*,¹⁵ samples are stored for several months to ensure that all traces of $PbCl_2$ (which is insoluble in apolar solvents) have precipitated. Hereafter, decantation separates the $PbCl_2$ from the suspension containing the Qdots, and a powder X-ray diffractogram is measured to determine whether $PbCl_2$ is still present.

In addition, 1H NMR is used to detect possible traces of OLA-S and/or TOPS complexes. Samples are prepared by drying a toluene suspension of PbS Qdots, followed by resuspension in deuterated toluene (toluene- d_6). 1H NMR spectra are measured at 295 K, using a Bruker DRX 500 spectrometer equipped with a TBI Z-gradient probe head. Absolute concentrations of all organic species in the suspensions are determined by adding a known amount (2 μ L) of CH_2Br_2 to the suspension and collecting the spectra under quantitative conditions.^{16,27}

Determination of the PbS Qdot Composition and Concentration. We calculate the concentration of PbS Qdots c_0 from the elemental concentrations of Pb and S, in combination with a calculation of the number of atoms N per Qdot of given d . First, a known amount of four PbS Qdot samples is digested in HNO_3 . The Pb concentration is then determined with ICP-MS. The measurements are performed with a PerkinElmer SCIEX Elan 5000 inductively coupled plasma mass spectrometer.

We are not able to accurately determine the S concentration with ICP-MS for several reasons. First, sulfur displays a high first ionization energy, resulting in a low sensitivity of Ar-based ICP-MS for this element. Second, signals arising from O_2^+ and HO_2^+ molecular ions are isobaric with signals arising from S ions and further hamper an accurate determination. Third, the preliminary digestion of PbS Qdot samples using concentrated HNO_3 im-

plies the risk of losing sulfur from the solution due to the formation of volatile H_2S . Therefore, we use RBS^{28,29} to measure the Pb/S atomic ratio R of five independent samples. As PbS Qdots are synthesized from $PbCl_2$, the Cl/Pb ratio is determined, as well. RBS samples are prepared by dropcasting a small amount of PbS Qdots on a MgO substrate, hereby forming a 100–200 nm close-packed thin film. The measurements are performed with a 2.5 MeV He^+ ion beam and two solid-state detectors placed at a backscattering angle of 165 and 167.6°, respectively. The high energy of the beam ensures that S and Cl, atoms with a comparable mass, provide resolved signals. From the backscattered yield, R is determined:

$$R = \frac{A_{Pb}}{A_S} \left(\frac{Z_S}{Z_{Pb}} \right)^2 \quad (17)$$

A_{Pb} denotes the area of the peak resulting from backscattering at Pb atoms, A_S equals the area of the S peak, Z_{Pb} and Z_S are the atomic number of Pb and S, respectively. A similar expression yields the Cl/Pb ratio. As both RBS detectors yield similar results, we report the average value.

Absorbance Measurements. To determine the relation between the Qdot TEM size d and the band gap E_0 , small amounts of the PbS Qdots used for the TEM measurements are dried and resuspended in 1 mL of C_2Cl_4 . The absorbance spectra are measured with a Perkin-Elmer Lambda 950 UV–vis–NIR spectrophotometer. We use a black walled self-masking microcell with a path length of $l = 1$ cm. E_0 is determined from the spectral position of the first absorption peak.

For the determination of ϵ , identical amounts of PbS Qdots as used for the ICP-MS measurements are dried and resuspended in 1 mL of C_2Cl_4 . From the absorbance A and c_0 , obtained from ICP-MS and RBS data, ϵ is determined using Beer's law: $A = \epsilon c_0 l$.

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